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Acids and Bases from a Different View.

In experiment 4, you learned about Bronsted Acid and Bases (looking ONLY at **oxi-acids/bases**)

- Bronsted Acid: proton donor
- Bronsted Base: proton acceptor

So what happens when an acid doesn't have any hydrogens (such as metal ions (in salt form) that will be used in this experiment?

Is it still an acid?

Here is an example that demonstrates whether an acid with no hydrogens remains as an acid!

[http://www.youtube.com/watch?v=3y5Lhk8XeJs&feature=player_embedded]

For more information:

Introduction of Lewis Acid/Base and an explanation of what is occurring in the case of AlCl_3 : <http://www.youtube.com/watch?v=lhbYoVFhAww&feature=related>

A change in the pH is observed, what is causing the change to occur?

What would the outcome be when NaCl is substituted?

[http://www.youtube.com/watch?v=hDOHuT35k9A&feature=player_embedded]