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Preparing a solution of known concentration

**Molarity**

Molarity is a measurement of concentration.

Specifically for molarity, it is the number of moles in a given volume

\[
\text{Molarity} = \frac{\text{moles}}{\text{liters}}
\]

There are 6 moles of NaCl in 3 liters of water, so the Molarity (Concentration) is 2 moles / liter. The molarity of the solution is 2.0

[http://www.youtube.com/watch?v=kuD2bDHVP90 &feature=player_embedded]
The main equation for calculating molarity is that 
molarity = the number of moles in one liter of 
solution

The video below shows exactly how to setup and 
use the molarity equation to determine the 
number of moles needed to make 100mL of a 
0.1M solution

[http://www.umich.edu/~chem125/softchalk/Exp 
2_Final/]

Now use the equation in the video to solve these 
problems. You may need to determine the 
molecular weight of compounds as well, so have 
your periodic table 
[http://www.webelements.com/] ready!

Still wanting some extra practice on calculating 
molarities, and volumes and moles? Visit the link 
below for a bottomless molarity worksheet!